



CHEMISTRY

CH:8 REDOX REACTIONS

Class: XI

1. What is meant by disproportionation reaction? Give an example
2. Define oxidation number.
3. Define oxidation number.
4. What is electrode potential?
5. What is electrochemical series?
6. Reduction potential of zinc is -0.76. What does it mean?
7. Answer the following questions based on the chemical reaction.
$$\text{Zn} + 2 \text{Ag}^+ \longrightarrow \text{Zn}^{2+} + 2 \text{Ag}$$
 taking place in a galvanic cell.
 - i) Write the reactions occurring at each electrode.
 - ii) Name the salt to be taken in salt bridge.
 - iii) In which direction do the electrons flow in the external circuit.
8. Determine the oxidation number of bromine in the following species.
i) $\underline{\text{K}}\text{Cl}$ ii) $\underline{\text{Br}}\text{O}_4^-$ iii) $\underline{\text{P}}\text{O}_4^{3-}$ iv) $\text{H}_2\underline{\text{S}}$
9. A layer of CuO on copper vessel can be easily cleaned by dilute HCl but oxide layer of aluminium on the aluminium vessel cannot be cleaned by HCl. Why?
10. Balance the following redox equations.
 - i) $\text{Br}_2 + \text{H}_2\text{O}_2 \longrightarrow \text{BrO}_3^- + \text{H}_2\text{O}$ in acidic medium
 - ii) $\text{As}_2\text{O}_3^{2-} + \text{IO}_3^- \longrightarrow \text{AsO}_4^{3-} + \text{I}^-$ in acidic medium